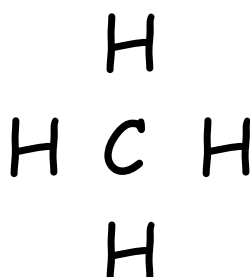


General Rules for Drawing Covalently Bonded Lewis Structures

1. The central atom most of the time is the more positive atom.
2. All atoms want to fulfill the octet rule.
Exceptions: Hydrogen 2 valence
Boron 6 valence
Aluminum 6 valence
Gallium 6 valence
Indium 6 valence
Sometimes Exceptions:
Nitrogen 5 valence
Phosphorus 10 valence
Sulfur 12 valence
3. Double and triple bonds are represented by shared two or three pairs of electrons.
4. Draw electrons in pairs when making a complete Lewis structure; nitrogen can be an exception
5. Look at the examples on this paper to guide you.

Example 1 CH₄



Example 2 H_2

H H

Example 3 CO_2

O C O

Example 3 CO

C O

Example 4 N_2

N N

Example 5 CH_3CHCH_2

