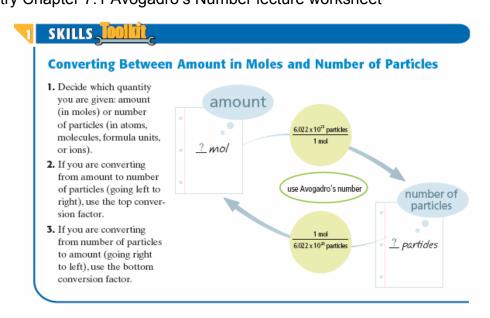
Name			_			
Regents Chemistry	Chapter 7.1	Avogadro's	Number	lecture	workshe	eet



Questions from the PowerPoint lecture:

Using the information provided on Slide #3,

a)	1	mo	e o	f any	' subs	tance	equals	how	many	of th	at sı	ubst	anc	e?	·
															_

e)	1 mole of sodium chloride con	ntains how many	NaCl formula units?	

a.	How many	/ Na⁺	ions?	

f۱	1 mole of sodium dichrom	ate contains how many	v K₀Cr₀O₂ formula units?	
ι,	i illole di sodidili dicilidi	iale contains now many	y N2O12O7 Iorriula urilla:	

a.	How many	/ K⁺ ions?		
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In Summary:

1 mole of sucrose $C_{12}H_{22}O_{11}$ contains 6.022 x 10^{23} molecules.

1 mole of KBr contains ___ mole of KBr formula units, ___ mole of K⁺ ions and ___ mole of Br⁻ ions

1 mole of CaCl₂ contains ____ mole of CaCl₂ formula units, ____ mole of Ca⁺² ions and ____ moles of Cl⁻ ions.

What's the difference between C₁₂H₂₂O₁₁, CaCl₂ and KBr?

Practice

1. How many ions are there in 0.187 mol of Na⁺ ions?

2. How many atoms are there in 1.45×10^{-17} mol of arsenic?

3. How many moles of xenon do 5.66×10^{23} atoms equal?

4. How many moles do 5.66×1025 lithium ions, Li⁺, equal?