Chapter 4.4 Reduction-Oxidation

1. Determine the oxidation number of each element in the following substances.

(a) KI

(e) As₄

(i) $C_r O_4^{2-}$

(b) CoCl₂

(f) K₂O₂

(j) $Ba(OH)_2$

(c) MnO_4

(g) Ga(OH)₃

(k) $Ca(NO_3)_2$

(d) HBrO

(h) HPO₄²⁻

2. Which element is oxidized an which element is reduced in the following equations. Write the Net ionic equation for each reaction.

1. $HCI(aq) + Mg(s) \rightarrow$

2. $Zn(NO_3)_2(aq) + Al(s) \rightarrow$

3. $Zn(s) + HCI(aq) \rightarrow$

4. $Pb(s) + AgNO_3(aq) \rightarrow$