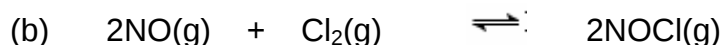
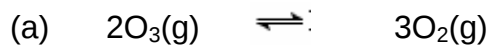


Section 15.1 – 15.3

1. Write the equilibrium expression for K_c for the following reactions:



2. Write the equilibrium-constant expression for $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$

3. In the synthesis of ammonia from nitrogen and hydrogen, $K_c = 9.60$ at 300°C :



Calculate K_p for this reaction at this temperature

4. For the equilibrium $2\text{SO}_3(\text{g}) \rightleftharpoons 2\text{SO}_2(\text{g}) + \text{O}_2(\text{g})$ at a temperature of 1000 K, K_c has the value of 4.08×10^{-3} . Calculate the value for K_p .

5. Write the equilibrium-constant expressions for K_c and K_p for each of the following reactions.

