

Name _____

Regents Chemistry

Naming Type III (covalent) compounds

Directions: Name the following covalent compounds using the following rules**Rules for naming type III compounds**

- 1. The first element in the formula is named first, and the full element name is used
- 2. The second element is named as though it were an anion (*-ide ending*)
- 3. Prefixes are used to denote the numbers of atoms present.
- 4. The prefix mono is *never* used for naming the first element

Table of Prefixe

Prefix	Number Indicated
<i>mono-</i>	1
<i>di-</i>	2
<i>Tri-</i>	3
<i>tetra-</i>	4
<i>penta-</i>	5
<i>hexa-</i>	6
<i>hepta-</i>	7
<i>octa-</i>	8

Compound	First element name	Prefix for first element	Second element name	Prefix for second element	Compound name
NO ₂					
PCl ₅					
P ₄ O ₆					
SF ₆					
SO ₃					
SO ₂					
N ₂ O ₃					
CO					
CO ₂					
SiO ₂					
O ₂ F ₂					
XeF ₆					
CBr ₄					

Directions: Write for formulas for the following compounds

Compound	Chemical Formula
iodine monochloride	
nitrogen trichloride	
silicon dioxide	
germanium tetrahydride	
dinitrogen tetrabromide	
diphosphorous pentasulfide	
Selenium dioxide	
Silicon dioxide	
dihydrogen oxide	
nitrogen trihydride	